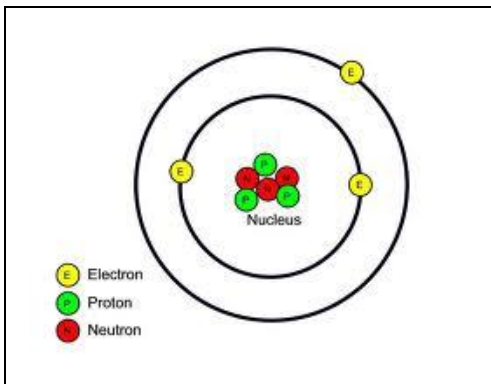


Name: _____ Class: _____ Date: _____

Atomic Structure Worksheet



**Protons, Neutrons,
and Electrons!**

Oh, My!

By Vicki – The Science Lady

Name: _____ Class: _____ Date: _____

Atomic Structure

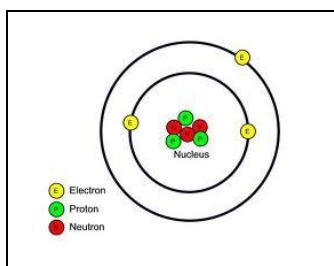
Use a periodic table to help you fill in the chart below. Mass numbers are rounded to the nearest whole number.

Remember:

Atomic number = number of protons = number of electrons

Mass number = number of protons + number of neutrons

Element	Symbol	Atomic Number	Mass Number	Number of Protons	Number of Neutrons	Number of Electrons
Beryllium		4				
	Ca		40			
				26		26
Silicon			28			
		18	40			
					10	9
Silver						
				54		



Name: _____ Class: _____ Date: _____

Atomic Structure **Answer Key**

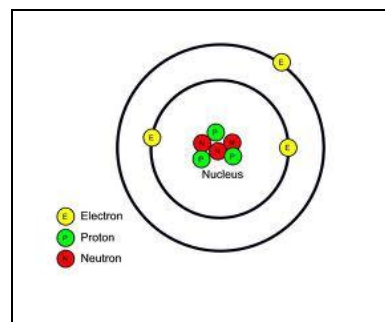
Use a periodic table to help you fill in the chart below. Mass numbers are rounded to the nearest whole number.

Remember:

Atomic number = number of protons = number of electrons

Mass number = number of protons + number of neutrons

Element	Symbol	Atomic Number	Mass Number	Number of Protons	Number of Neutrons	Number of Electrons
Beryllium	Be	4	9	4	5	4
Calcium	Ca	20	40	20	20	20
Iron	Fe	26	56	26	30	26
Silicon	Si	14	28	14	14	14
Argon	Ar	18	40	18	22	18
Fluorine	F	9	19	9	10	9
Silver	Ag	47	108	47	61	47
Xenon	Xe	54	131	54	77	54



Name: _____ Class: _____ Date: _____

Teacher Notes

Students can use a separate periodic table, or use one from their textbook.

This review can teach them how to find their way around the periodic table.

It should be done after they have been introduced to atomic structure. Remind the students that the atomic number is equal to the number of protons, which is equal to the number of electrons, and it written at the top of the sheet. Atomic mass is equal to the number of protons plus the number of neutrons, so they can find the number of neutrons by subtracting the number of protons from the atomic mass.

I use this as a quick review to reinforce basic facts that can be found by using the periodic table. It could be used as homework if they have access to a periodic table, or incorporate it in a quiz or test. You might want to 'white-out' the give-aways: atomic number = number of protons, etc, if you use it in a test, but be sure they have a periodic table to use.